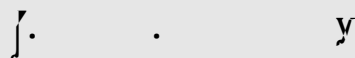




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$$T = (278.15 \quad 333.15)$$

Y, *, W, W, W, Y

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ABSTRACT

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 (y + y)
 278.15 333.15
 (278.15 333.15),
 A y B y λ y.
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1. Introduction

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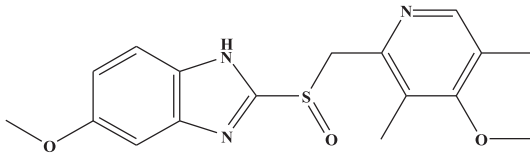


FIGURE 1.

2. Experimental

2.1. Materials and apparatus

The purity of the materials was determined by gas chromatography-mass spectrometry (GC-MS) and was found to be ≥ 0.980 . The molecular weight (M) of the compound is 244.32. The boiling point (BP) of the compound is 180 °C at 0.1 mmHg. The density (d) of the compound is 1.25 g/cm³ at 20 °C. The refractive index (n_D²⁰) of the compound is 1.5120. The compound was stored in a dark, dry, and cool place. The compound was used without further purification.

2.2. Methods

The vapor pressure (p) of the compound was measured by the static method. The temperature (T) of the measurement was 431.15, 430.65, 428.15, and 427.65 K. The vapor pressure (p) of the compound was measured by the static method. The temperature (T) of the measurement was 431.15, 430.65, 428.15, and 427.65 K. The vapor pressure (p) of the compound was measured by the static method. The temperature (T) of the measurement was 431.15, 430.65, 428.15, and 427.65 K.

TABLE 1

M	M	M	M
W	345.42	0.980	M
M	18.02	0.995	M
	32.04	0.997	M
1-B	46.07	0.990	M
A	74.12	0.999	M
A	41.05	0.995	M
y	58.08	0.997	M
y	88.11	0.997	M
	72.11	0.990	M

The vapor pressure (p) of the compound was measured by the static method. The temperature (T) of the measurement was 431.15, 430.65, 428.15, and 427.65 K. The vapor pressure (p) of the compound was measured by the static method. The temperature (T) of the measurement was 431.15, 430.65, 428.15, and 427.65 K.

3. Results and discussion

3.1. In pure solvents

3.1.1. Solubility and correlation models

The solubility (x) of the compound in pure solvents was measured by the static method. The temperature (T) of the measurement was 278.15, 333.15, and 333.15 K. The solubility (x) of the compound in pure solvents was measured by the static method. The temperature (T) of the measurement was 278.15, 333.15, and 333.15 K.

$$RD = \frac{x_i - x_{ci}}{x_i} \tag{3}$$

The relative deviation (RD) of the compound in pure solvents was calculated by the static method. The temperature (T) of the measurement was 278.15, 333.15, and 333.15 K. The relative deviation (RD) of the compound in pure solvents was calculated by the static method. The temperature (T) of the measurement was 278.15, 333.15, and 333.15 K.

$$A = \frac{1}{N} \sum_{i=1}^N \frac{x_i - x_{ci}}{x_i} \tag{4}$$

TABLE 2

T/ M	$\gamma(x)$	(278.15 333.15)		0.1M ^{a,b}	
		1000x	100	. (5)	. (6)
<i>Water</i>					
278.15	0.0407		6.14	-9.15	-5.35
283.15	0.0500		2.03	-8.45	-5.81
288.15	0.0615		-0.69	-7.00	-5.39
293.15	0.0764		-1.02	-3.89	-3.14
298.15	0.0936		-1.46	-1.75	-1.69
303.15	0.1145		-0.97	0.52	0.09
308.15	0.1377		-0.95	1.58	0.84
313.15	0.1652		-0.13	2.75	1.87
318.15	0.1959		0.58	3.13	2.33
323.15	0.2297		1.09	2.68	2.15
328.15	0.2631		0.16	0.14	0.11
333.15	0.2994		-0.53	-2.89	-2.11
<i>Methanol</i>					
278.15	2.058		3.33	-9.66	-7.27
283.15	2.639		0.41	-8.70	-6.98
288.15	3.480		1.65	-3.95	-2.87
293.15	4.350		-1.10	-4.10	-3.50
298.15	5.691		1.97	1.08	1.26
303.15	6.863		-1.90	-1.33	-1.46
308.15	8.573		-1.10	0.39	0.05
313.15	10.55		-0.69	1.21	0.75
318.15	12.89		-0.06	1.73	1.29
323.15	15.62		0.81	2.01	1.69
328.15	18.61		0.81	0.95	0.90
333.15	21.65		-0.62	-2.06	-1.67
<i>Ethanol</i>					
278.15	1.574		7.74	0.37	3.55
283.15	1.875		2.51	-2.76	-0.43
288.15	2.243		-1.56	-4.89	-3.37
293.15	2.612		-7.61	-9.30	-8.50
298.15	3.465		0.81	0.56	0.70
303.15	4.130		-0.90	-0.20	-0.50
308.15	5.085		1.45	2.70	2.12
313.15	6.047		1.12	2.60	1.87
318.15	7.019		-0.91	0.47	-0.23
323.15	8.316		-0.19	0.70	0.22
328.15	9.818		0.84	0.87	0.83
333.15	11.26		-0.44	-1.66	-1.02
<i>1-Butanol</i>					
278.15	1.282		6.99	-5.84	-2.64
283.15	1.532		-0.36	-9.60	-7.25
288.15	1.962		0.30	-5.21	-3.78
293.15	2.404		-2.17	-4.88	-4.15
298.15	3.026		-0.71	-1.20	-1.06
303.15	3.702		-0.96	0.08	-0.21
308.15	4.550		0.31	2.25	1.68
313.15	5.401		-0.89	1.43	0.72
318.15	6.504		0.33	2.42	1.75
323.15	7.788		1.86	3.19	2.74
328.15	8.842		-1.03	-0.98	-1.02
333.15	10.35		-0.02	-1.85	-1.23
<i>Acetonitrile</i>					
278.15	1.177		5.78	-6.29	-3.07
283.15	1.425		-0.26	-8.87	-6.52
288.15	1.821		0.20	-4.97	-3.52
293.15	2.221		-2.80	-5.39	-4.64
298.15	2.732		-3.77	-4.28	-4.12
303.15	3.538		1.57	2.49	2.22
308.15	4.286		1.27	3.04	2.49
313.15	5.126		0.65	2.77	2.09
318.15	6.075		0.05	2.01	1.35
323.15	7.113		-0.88	0.40	-0.06
328.15	8.431		0.27	0.32	0.28
333.15	9.775		0.00	-1.69	-1.09
<i>Acetone</i>					
278.15	1.463		4.30	-4.47	-0.95
283.15	1.731		-0.89	-7.03	-4.52
288.15	2.155		-0.09	-3.67	-2.17
293.15	2.648		0.35	-1.23	-0.54

(continued on next page)

TABLE 2 (continued)

T/	1000x	100		
		(5)	(6)	(9)
298.15	3.126	-2.25	-2.33	-2.27
303.15	3.773	-1.76	-0.75	-1.18
308.15	4.638	1.41	3.01	2.30
313.15	5.415	0.22	2.05	1.21
318.15	6.372	0.56	2.19	1.41
323.15	7.388	0.13	1.15	0.64
328.15	8.496	-0.45	-0.48	-0.49
333.15	9.815	0.07	-1.47	-0.71
Ethyl acetate				
278.15	0.9807	3.66	2.86	6.65
283.15	1.124	-0.27	-0.98	1.76
288.15	1.338	0.02	-0.52	1.15
293.15	1.565	-1.03	-1.38	-0.60
298.15	1.854	-0.45	-0.59	-0.54
303.15	2.199	0.66	0.72	0.22
308.15	2.535	-0.74	-0.52	-1.37
313.15	2.927	-1.62	-1.29	-2.29
318.15	3.426	-0.75	-0.41	-1.31
323.15	4.049	1.36	1.62	1.05
328.15	4.696	1.92	1.96	1.97
333.15	5.226	-1.37	-1.71	-0.83
Tetrahydrofuran				
278.15	4.577	12.68	-6.92	-10.04
283.15	6.437	10.74	-2.96	-5.29
288.15	8.434	4.13	-5.27	-7.03
293.15	11.31	1.43	-4.07	-5.26
298.15	14.62	-3.18	-5.70	-6.42
303.15	19.63	-2.13	-2.37	-2.68
308.15	25.88	-1.20	-0.04	-0.03
313.15	32.73	-2.88	-0.99	-0.77
318.15	42.45	-0.46	1.42	1.74
323.15	54.51	2.34	3.65	3.93
328.15	66.34	1.19	1.46	1.60
333.15	78.86	-1.06	-2.35	-2.53

^a $u_i(x) = 2.00\%$ y $u(T) = 0.05$, $u(p) = 2$.

3.1.2. The modified Apelblat model

$$x = A + \frac{B}{(T/K)} + C \quad (7)$$

where A, B, C are parameters to be determined. The parameters A, B, C are determined by fitting the experimental data. The parameters A, B, C are determined by fitting the experimental data. The parameters A, B, C are determined by fitting the experimental data.

3.1.3. Buchowski Ksiazaczak λ h model

$$1 + \frac{\lambda(1-x)}{x} = \lambda h \left[\frac{1}{(T/K)} - \frac{1}{(T_m/K)} \right] \quad (6)$$

where λ and h are parameters to be determined.

3.1.4. Ideal model

$$x\gamma = \frac{\Delta_{diss}H}{R} \left[\frac{1}{T_m} - \frac{1}{T} \right] \quad (7)$$

$$A^{def} = \frac{\Delta_{diss}H}{R} \times \frac{1}{T_m}, \quad B^{def} = -\frac{\Delta_{diss}H}{R} \quad (8)$$

$$x = A + \frac{B}{T} \quad (9)$$

3.1.5. Thermodynamic parameters

$$\Delta H_m^0 = -R \times \frac{x}{(1/T)} \quad (10)$$

where ΔH_m^0 is the enthalpy of mixing at the ideal state.

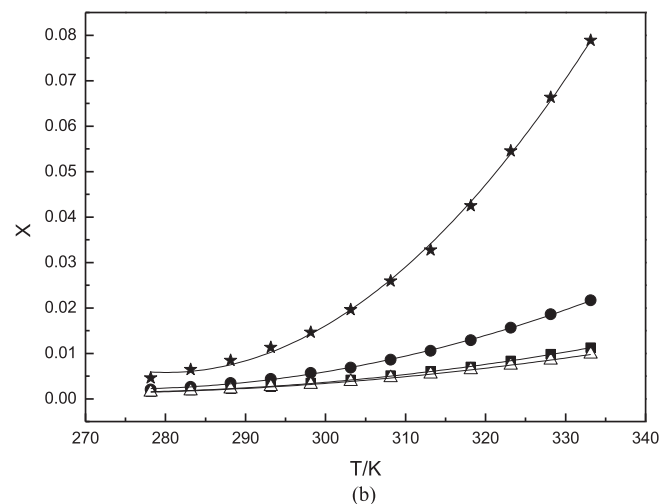
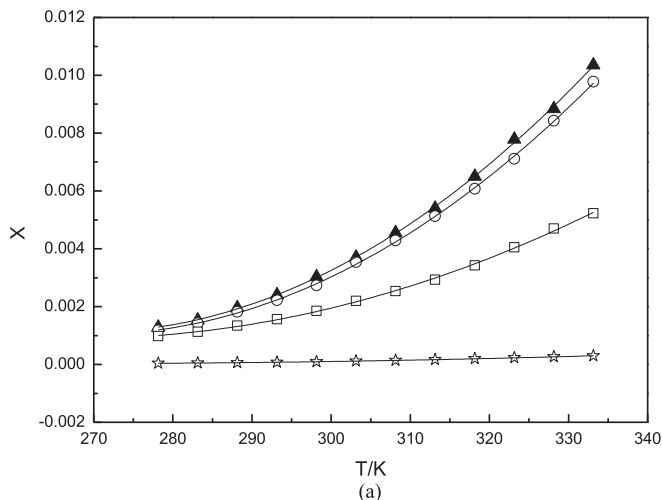


FIGURE 2. ΔH_m^0 (a) and ΔG_m^0 (b) vs. T/K . Symbols: (▲) 1-B; (○) A; (□) y; (☆) W; (●) M; (★) y.

$$\Delta H_m^0 = -(b - cT)R \quad (11)$$

$$\Delta G_m^0 = -RT \ln \left(\frac{p}{p^0} \right) \quad (12)$$

TABLE 3

	A	B	C	$10^2 A$
W	128.17	-9170.28	-18.72	1.31
M	110.16	-8641.07	-15.16	1.20
	53.88	-5627.47	-7.14	2.17
1-B	110.12	-8334.80	-15.44	1.33
A	100.63	-7929.87	-14.02	1.46
A	65.40	-6002.12	-8.95	1.04
y	-34.21	-1199.77	5.61	1.15
y	207.2	-13842.9	-28.9	3.62
				$(10^2 A) = 13.29$

TABLE 4

	100λ	$10^2 A$	$10^2 A$
W	0.21	1453065	3.66
M	25.80	14540.97	3.10
	8.66	36768.28	2.26
1-B	8.72	37881.21	3.24
A	8.47	39469.21	3.54
A	6.36	46625.85	2.49
y	2.64	100547.6	1.21
y	210.30	2278.137	3.10
		$(10^2 A) = 22.60$	

TABLE 5

	A	B	$10^2 A$
W	1.84	-3310.50	2.57
M	7.82	-3875.08	2.47
	5.71	-3392.12	1.95
1-B	5.93	-3495.80	2.35
A	5.99	-3533.63	2.62
A	5.00	-3203.27	1.53
y	3.61	-2948.77	1.15
y	11.52	-4676.96	3.94
		$(10^2 A) = 18.60$	

$$\Delta S_m^0 = \frac{\Delta H_m^0 - \Delta G_m^0}{T} \quad (13)$$

$$\% \xi_H = \frac{|\Delta H_m^0|}{|\Delta H_m^0| + |T \cdot \Delta S_m^0|} \times 100\% \quad (14)$$

$$\% \xi_S = \frac{|T \cdot \Delta S_m^0|}{|\Delta H_m^0| + |T \cdot \Delta S_m^0|} \times 100\% \quad (15)$$

278.15 333.15 $\Delta H_m^0, \Delta S_m^0, \Delta G_m^0, \% \xi_H, \% \xi_S$ 6.

3.2. Binary solvent mixtures

3.2.1. Solubility and correlation models

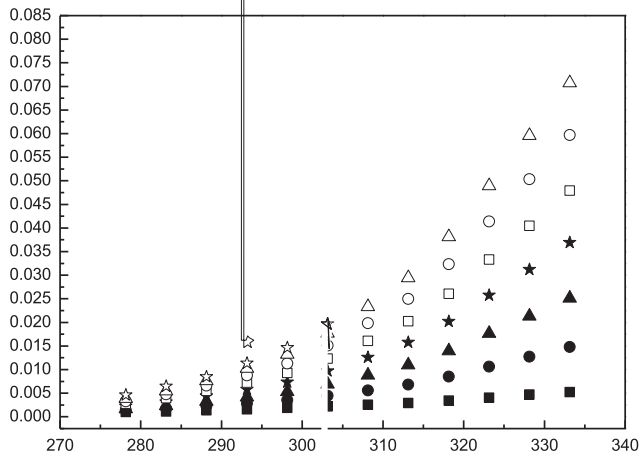
$$\ln \left(\frac{p}{p^0} \right) = \ln \left(\frac{p^*}{p^0} \right) + \ln \left(\frac{p}{p^*} \right)$$

TABLE 6

	ΔH_m^0	ΔS_m^0	ΔG_m^0	$\% \xi_H$	$\% \xi_S$
W	28.67	19.49	22.87	83.84	16.16
M	33.32	68.88	12.39	61.40	38.60
	28.64	49.04	13.71	65.72	34.28
1-B	30.06	52.93	14.01	65.21	34.79
A	30.30	53.10	14.18	65.28	34.72
A	27.16	43.47	13.94	67.26	32.74
y	24.23	28.93	15.34	73.30	26.70
y	41.65	105.44	9.66	56.51	43.49

TABLE 7
M.

x_A	$y(x)$	$(y + y)$	y	278.15	333.15	$0.1M_{a,b}$
	100x	100x - x	$/x(. (5))$	100x - x	$/x(. (17))$	100x - x
						$/x(. (20))$
<i>T</i> = 278.15 K						
0.0000	0.0981	3.6649		0.0000		5.5665
0.1748	0.1348	5.9638		5.5327		1.8428
0.3397	0.1652	1.2125		4.1823		1.5203
0.4955	0.2227	5.2449		0.1319		1.9871
0.6429	0.2767	6.7786		1.2538		2.1144
0.7826	0.3242	5.5088		0.0160		2.2598
0.9153	0.3882	8.9930		0.5848		1.7342
1.0000	0.4577	12.6848		0.0000		0.7236
<i>T</i> = 283.15 K						
0.0000	0.1124	0.2684		0.0000		4.8445
0.1748	0.1714	2.3923		3.5859		1.5088
0.3397	0.2358	4.2893		0.5303		0.6774
0.4955	0.3108	4.0902		1.5784		1.3937
0.6429	0.3905	5.7810		0.1742		1.4349
0.7826	0.4756	7.3272		2.1831		1.1860
0.9153	0.5453	6.3833		1.9287		1.2373
1.0000	0.6437	10.7355		0.0000		0.2921
<i>T</i> = 288.15 K						
0.0000	0.1338	0.0192		0.0000		8.5736
0.1748	0.2260	3.4979		0.9541		3.1181
0.3397	0.3254	5.7285		0.7396		0.2739
0.4955	0.4389	5.9433		0.3892		1.1966
0.6429	0.5454	5.7368		0.3353		1.5966
0.7826	0.6589	5.8666		0.3530		1.3970
0.9153	0.7654	5.7591		0.2405		1.3186
1.0000	0.8434	4.1327		0.0000		0.9861
<i>T</i> = 293.15 K						
0.0000	0.1565	1.0316		0.0000		9.0428
0.1748	0.2832	0.7754		1.2106		3.1296
0.3397	0.4197	2.3069		0.8292		0.5059
0.4955	0.5756	2.5284		0.3702		1.2927
0.6429	0.7218	2.4412		0.3262		1.6508
0.7826	0.8778	2.5783		0.3813		1.3477
0.9153	1.024	2.5414		0.2698		1.2624
1.0000	1.131	1.4290		0.0000		0.8699
<i>T</i> = 298.15 K						
0.0000	0.1854	0.4524		0.0000		7.5555
0.1748	0.3514	1.9198		1.4146		2.3090
0.3397	0.5301	1.7739		0.8929		0.9857
0.4955	0.7344	2.0184		0.3503		1.5828
0.6429	0.9259	2.2710		0.3160		1.9021
0.7826	1.130	2.3162		0.4016		1.6145
0.9153	1.322	2.4211		0.2918		1.5817
1.0000	1.462	3.1847		0.0000		1.2283
<i>T</i> = 303.15 K						
0.0000	0.2199	0.6572		0.0000		12.2283
0.1748	0.4465	1.1276		1.7652		4.4016
0.3397	0.6905	1.2900		0.9918		0.9167
0.4955	0.9694	1.4826		0.3114		1.5215
0.6429	1.231	1.6392		0.2955		1.9652
0.7826	1.510	1.6802		0.4334		1.3696
0.9153	1.771	1.7388		0.3275		1.3651
1.0000	1.963	2.1316		0.0000		0.7413
<i>T</i> = 308.15 K						
0.0000	0.2535	0.7413		0.0000		8.2666
0.1748	0.5570	1.2135		2.2370		3.7023
0.3397	0.8838	1.1731		1.1098		0.3657
0.4955	1.257	1.1371		0.2545		0.4095
0.6429	1.607	1.1322		0.2656		0.6697
0.7826	1.981	1.1100		0.4722		0.0570
0.9153	2.331	1.0983		0.3720		0.1113
1.0000	2.588	1.2049		0.0000		0.4408
<i>T</i> = 313.15 K						
0.0000	0.2927	1.6211		0.0000		8.6342
0.1748	0.6801	2.5013		2.6160		4.6684
0.3397	1.097	2.7898		1.1954		1.9404
0.4955	1.574	2.8591		0.2072		1.4026
0.6429	2.021	2.8849		0.2411		1.8411
0.7826	2.498	2.9148		0.5011		0.7072
0.9153	2.945	2.9194		0.4055		1.0413
1.0000	3.273	2.8761		0.0000		0.1045



ΔH_m^0 ΔG_m^0 $\% \xi_H \geq 56.51\%$

4. Conclusions

W 2 11 2 3:(1) 278.15 333.15 y

;(2)

(, A A , λ) , -

y , B / \int y A) y \int 15-642

B / M A A

y B / y A ;(3) y- , -

y.

Acknowledgments

y y y \int -

((14)2057), \int y y

(B 2014386).

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